

Chemistry Stoichiometry Problem Sheet 2 Answer Key

Chemistry Stoichiometry Problem Sheet 2 Answer Key Chemistry Stoichiometry Problem Sheet 2 Answer Key This answer key provides detailed solutions and explanations for the problems presented in Chemistry Stoichiometry Problem Sheet 2 The problems focus on key concepts in stoichiometry including Mole conversions Converting between grams moles and particles atoms molecules or formula units Balancing chemical equations Ensuring the number of atoms of each element is equal on both sides of the equation Stoichiometric calculations Using mole ratios from balanced equations to determine the amount of reactants or products involved in a reaction Limiting reactants Identifying the reactant that limits the amount of product that can be formed Percent yield Calculating the efficiency of a reaction based on the actual yield obtained compared to the theoretical yield

Problem 1 Mole Conversions

a How many moles are in 250 g of sodium chloride NaCl
Solution 1 Molar mass Find the molar mass of NaCl Na 2299 gmol Cl 3545 gmol 5844 gmol
2 Conversion Use the molar mass to convert grams to moles 250 g NaCl 1 mol NaCl 5844 g NaCl 0427 mol NaCl
b How many grams are in 0750 mol of carbon dioxide CO₂
Solution 1 Molar mass Find the molar mass of CO₂ C 1201 gmol 2O 1600 gmol 4401 gmol
2 Conversion Use the molar mass to convert moles to grams 2 0750 mol CO₂ 4401 g CO₂ 1 mol CO₂ 330 g CO₂
c How many particles are in 150 mol of water H₂O
Solution 1 Avogadros number Use Avogadros number 6022 x 10²³ particlesmol to convert moles to particles 150 mol H₂O 6022 x 10²³ molecules H₂O 1 mol H₂O 903 x 10²³ molecules H₂O

Problem 2 Balancing Chemical Equations

Balance the following chemical equation CH₄ + O₂ → CO₂ + H₂O
Solution 1 Count atoms On the reactant side we have 1 C 4 H and 2 O On the product side we have 1 C 2 H and 3 O
2 Adjust coefficients To balance the equation we need to adjust the coefficients in front of each molecule CH₄ + 2O₂ → CO₂ + 2H₂O
3 Recount atoms Now we have 1 C 4 H and 4 O on both sides of the equation

Problem 3 Stoichiometric Calculations

Consider the balanced equation 2Na + Cl₂ → 2NaCl

a How many grams of sodium chloride NaCl can be produced from 500 g of sodium Na
Solution 1 Molar mass Find the molar masses of Na 2299 gmol and NaCl 5844 gmol
2 Convert grams to moles 500 g Na 1 mol Na 2299 g Na 0217 mol Na
3 Use mole ratio The balanced equation shows a 22 mole ratio between Na and NaCl This means 0217 mol Na will produce 0217 mol NaCl
4 Convert moles to grams 0217 mol NaCl 5844 g NaCl 1 mol NaCl 127 g NaCl
b How many grams of chlorine gas Cl₂ are required to react completely with 100 g of sodium Na
Solution 1 Convert grams to moles 100 g Na 1 mol Na 2299 g Na 0435 mol Na
2 Use mole ratio The balanced equation shows a 21 mole ratio between Na and Cl₂ This 3 means 0435 mol Na will react with 0217 mol Cl₂
3 Convert moles to grams 0217 mol Cl₂ 7090 g Cl₂ 1 mol Cl₂ 154 g Cl₂

Problem 4 Limiting Reactants

Suppose 200 g of magnesium Mg reacts with 300 g of oxygen O₂ according to the following balanced equation 2Mg + O₂ → 2MgO

a Which reactant is the limiting reactant
Solution 1 Convert grams to moles 200 g Mg 1 mol Mg 2431 g Mg 0823 mol Mg 300 g O₂ 1 mol O₂ 3200 g O₂ 0938 mol O₂
2 Determine limiting reactant Mg to MgO ratio 2 moles Mg 2 moles MgO 11 O₂ to MgO ratio 1 mole O₂ 2 moles MgO
12 Calculate moles of MgO

produced 0.823 mol Mg, 2 mol MgO, 2 mol Mg, 0.823 mol MgO, 0.938 mol O₂, 2 mol MgO, 1 mol O₂, 1.876 mol MgO. The limiting reactant is Mg as it produces less MgO. b) How many grams of magnesium oxide (MgO) can be produced? Solution 1: Use limiting reactant. Since Mg is the limiting reactant, we use its moles (0.823 mol) to calculate MgO production. 2) Convert moles to grams: 0.823 mol MgO × 40.30 g/mol = 33.2 g MgO.

Problem 5: Percent Yield. In a reaction, 150 g of sodium chloride (NaCl) is theoretically expected. However, only 125 g of NaCl is actually produced. What is the percent yield? Solution 1: Percent yield formula: $\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$. Plug in values: $\text{Percent yield} = \frac{125 \text{ g}}{150 \text{ g}} \times 100 = 83.3\%$.

Conclusion: This answer key provides a comprehensive understanding of the concepts and calculations involved in stoichiometry. By working through these problems and understanding the 4 explanations, you can gain a strong foundation in solving stoichiometry problems in chemistry. Remember to always balance chemical equations before performing any stoichiometric calculations. Use mole ratios from balanced equations to determine the amounts of reactants and products involved. Identify the limiting reactant to determine the maximum amount of product that can be formed. Calculate percent yield to assess the efficiency of a reaction. Stoichiometry is a crucial aspect of chemistry, enabling us to predict and analyze chemical reactions. With practice and a solid understanding of the concepts, you can master this important skill.

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